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$\ln K = (\Delta_r S/R) - (\Delta_r H/RT)$ The equilibrium for the dissociation of one mole of iodine vapour is 3.94×10^{-4} at 900K and the standard enthalpy is 154 kJmol⁻¹. Calculate

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the equilibrium constant at 1200K

assuming the standard enthalpy remains constant. $\ln(K_2/K_1) = -\Delta_r H^\circ/R (1/T_2 - 1/T_1)$, ANS = 6.7×10^{-2} at 1200K.

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