

## Chemistry Gases Unit Study Guide Answers

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Chemistry Gases Unit Study Guide The ideal gas law is expressed by the formula PV = nRT where P = pressure V = volume n = number of moles of gas R = ideal gas constant T = absolute temperature The value of R depends on the units of pressure, volume and temperature. Chemistry Study Guide for Gases - ThoughtCo

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The study of gases allows us to understand the behavior of matter at its simplest: individual particles, acting independently, almost completely uncomplicated by interactions and interferences between each other.

[Unit 1: Gases – Chemistry LibreTexts](#)

Chemistry Gases Unit Study Guide A gas is a state of matter with no defined shape or volume. Gases have their own unique behavior depending on a variety of variables, such as temperature, pressure, and volume. While each gas is different, all gases act in a similar matter.

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Force exerted by particles per unit area What is the formula for gas pressure? P=F/A where P=pressure and F=force of the gas particles and A=area What are some units for pressure?

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☐☐ study guide 3.6 Deviations from the Ideal Gas Law In the real world, gases don't always behave as defined by the Kinetic Molecular Theory. Conditions of high pressure and low temperature will cause gases to deviate from ideal gas behavior for the following reasons:

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Unit 8 Study Guide Chemistry Answers - modapktown.com The mole is the most common unit used to express the quantity of a chemical substance. For all solids, liquids, and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the Page 9/14